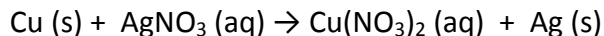


Mass Appeal HW

Read and outline (Cornell style) **Section 12.2** in your chemistry textbook. ****Your section outline must be at least ½ page and have 3-4 topic questions.**** Then answer the following assessment questions:

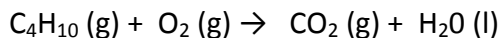
1. Why is a balanced chemical equation needed in solving stoichiometric calculations?

2. Consider the following reaction:



- Balance the equation
- What is the mole ratio of Cu to $\text{Cu(NO}_3)_2$?
- What is the mole ratio of AgNO_3 to $\text{Cu(NO}_3)_2$?
- What is the mole ratio of AgNO_3 to Ag?
- How many moles of Cu, copper, are needed to react with 6 moles of silver nitrate, AgNO_3 ?
- If 10 grams of Ag are produced, how many grams of AgNO_3 is used? Show your steps.

3. Consider the following equation:



- Balance the chemical equation.
- List the mole ratios of each reactant to each product. (There will be 4 ratios for your answer.)
- If 5 moles of ethane, C_4H_{10} are used, how many moles of water are produced?
- If 10 moles of oxygen, O_2 are used, how many moles of carbon dioxide are produced?
- How many grams of CO_2 will be produced if 42 grams of ethane, C_4H_{10} are used? Show your steps.